FROM SALT ANALYSIS-

* What is nesslers reagent?
* What is blue and green tinted ash?
* Study reagents for each cations.
* Why con.HCl is used in flame test?
* What is magneson reagent?
* Study the cause behind each precipitate of both identification and confirmatory test. Also study the chemical equations related to it
* What is the Brown precipitate formed in when nesslers reagent treated with sulphate ion.
* What is Brown ring?
* Why Paper ball is introduced in confirmatory test of nitrate?
* In ring test why do u add **freshly prepared Feso4?**
* Why ammonia is in 0 group?
* Explain hardy schulz rule
* Why is heating done for Oxalic Acid, KMnO4 titration?
* Why is H2SO4 used instead of HCl?
* Explain the principle of Salt analysis.
* What examples of salt with Colours?
* Why are salts classified into groups?
* What is Importance of Flame Test?
* Why is HCl used in Flame Test?
* Explain the role of Group 3 reagents.
* What is Common Ion effect?
* What is Formula of Potash Alum, Mohr’s Salt?

QUESTIONS FROM ORGANIC ANALYSIS—

* What is the identification test for phenol.
* Why violet color is obtained in neutral ferric chloride test?
* Distinguish between aldehyde and ketones.
* What is schiffs base?
* What is Rochelle salt?
* Equation of fehlling test and tollens test

QUESTIONS FROM TITRATION-

* What is Mohr salt?
* What is oxidizing agent in titration with KMnO4 and Mohr salt
* Oxidation state of Mn in KMnO4
* Why dil.H2SO4 added during titration
* indicator other than KMnO4
* Why titration done in acidic medium?
* How will you prepare M/20 mohr salt solution?
* How will you prepare M/40 mohr salt solution?
* What is molecular mass of KMnO4 ( i think i forget this question's proper form)
* To name and write some reagents(organic)

**What is the principle of volumetric analysis ?**  
**Ans.** In volumetric analysis, the concentration of a solution is determined by allowing a known volume of this to react quantitatively with another solution of known concentration.  
**7. What is titration ?**  
**Ans.** The process of adding one solution from the burette to another in the titration flask in order to complete the chemical reaction involved, is known as titration.  
**8. What is indicator ?**  
**Ans.** Indicator is a chemical substance which changes colour at the end point.  
**9. What is end point ?**  
**Ans.** The stage during titration at which the reaction is just complete is known as the end point of titration.  
**10. Why a titration flask should not be rinsed ?**  
**Ans.**This is because during rinsing some liquid will remain sticking to the titration flask there-fore the pipetted volume taken in the titration flask will increase.  
**11. What are primary and secondary standard substances ?**  
**Ans.**A substance is known as primary standard if it is- available in high degree of purity, if it is stable and unaffected by air, if it does not gain or lose moisture in air, if it is readily soluble and its solution in water remains as such for long time.On the other hand, a substance which does not possess the above characteristics is called a second-ary standard substance. Primary standards are crystalline oxalic acid, anhydrous Na2CO3, Mohr’s salt, etc.  
**12. Burette and pipette must be rinsed with the solution with which they are filled, why ?**  
**Ans.**The burette and pipette are rinsed with the solution with which they are filled in order to remove any substance sticking to their sides, which otherwise would decrease the volume of the liquids to be taken in them.  
**13. It is customary to read lower meniscus in case of colourless and transparent solutions and upper meniscus in case of highly coloured solutions, why ?**  
**Ans.** Because it is easy to read the lower meniscus in case of colourless solutions, while the upper meniscus in case of coloured solutions.  
**14. What is a normal solution ?**  
**Ans.**A normal solution is a solution, a litre of which contains one gm-equivalent of the solute. This is symbolised as 1 N.  
**15. Why the last drop of solution must not be blown out of a pipette ?**  
**Ans.** Since the drops left in the jet end is extra of the volume measured by the pipette.  
**16. Pipette should never be held from its bulb, why ?**  
**Ans.**The heat of our body may expand the glass bulb and introduce an error in the measurement of the volume.  
**17. What is acidimetry and alkalimetry ?**  
**Ans.** It is the branch of volumetric analysis involving chemical reaction between an acid and a base.  
**18. What do you mean by 1.0 M solution ?**  
**Ans.** A solution containing 1 mole of solute per litre of solution is 1.0 M solution.  
**19. What is meant by the term ‘concordant readings’ ?**  
**Ans.** The readings in volumetric analysis whcih differ by less than 0.05 mL Eire known as concordant readigns. .  
**20. Can one take oxalic acid solution in the burette and sodium hydroxide solution in the titration flask ? What are the limitations of doing so if any ?**  
**Ans.** No, because when sodium hydroxide solution is taken in the titration flask, the colour change at the end point would be pink to colourless. The accuracy in noting this change may be less as compound to change from colourless to pink.